Student Performance by Question
Core 1 / S-14

Chemistry (Multiple-Choice ONLY)

Teacher: _____

Analysis of Performance by Question of each student targeted for Remediation or Intervention
Enter a "1" in each box to indicate a missed question.

	Scientific Investigation										Atomic Structure and Periodic Relationships								Chemical Formulas and Reactions										
Student Name	Calculate the percent error in the results of a controlled experiment.	Identify the correct equipment to collect the results of a safe controlled experiment.	Calculate results of a controlled experiment using significant digits conventions.	Recognize and follow safe and appropriate laboratory procedures.	Recognize and follow safe and appropriate laboratory procedures.	Design safe controlled investigations.	Interpret results using appropriate error analysis.	Recognize and follow safe and appropriate laboratory procedures.	Identify variables in laboratory procedures.	Calculate the percent error in the results of a controlled experiment.	Apply information provided by the periodic table to calculate oxidation number.	Recognize that properties can be determined from the periodic table.	Evaluate information derived from historical models to make conclusions about matter.	Recognize that periodic trends can be determined from the periodic table.	Recognize that properties can be determined from the periodic table.	Recognize that properties can be determined from the periodic table.	Apply information provided by the periodic table to construct models.	Evaluate information to calculate properties of isotopes.	Classify chemical equations.	Classify chemical equations.	Describe compounds and bonding to provide appropriate chemical names.	Describe compounds and bonding to provide appropriate chemical names.	Interpret factors that affect equilibrium and kinetics.	Describe bonding to provide appropriate properties of substances.	Balance chemical equations.	Describe compounds and bonding to provide appropriate chemical names.	Interpret factors that affect equilibrium and kinetics.	Balance chemical equations.	Describe bonding to provide appropriate properties of substances.
Total	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0

					Molar Relationships								Phases of Matter and Kinetic Molecular Theory								
Identify the empiracal formula of a substance.	Use bonding principles to explain the interaction of substances.	Describe compounds and bonding to provide appropriate chemical names.	Classify chemical equations.	Describe bonding to provide appropriate properties of substances.	Apply molar relationships to perform calculations involving molar concentrations.	Apply molar relationships to perform calculations involving molar concentrations.	Interpret factors that affect solution chemistry.	Apply molar relationships to perform calculations involving molar stoichiometry.	Apply molar relationships to perform calculations involving molar concentrations.	Apply molar relationships to perform calculations involving molar concentrations.	Apply molar relationships to perform calculations involving molar stoichiometry.	Interpret factors that affect solution chemistry.	Apply molar relationships to perform calculations involving molar conversions.	Predict the behavior of matter and interactions between particles in colligative properties.	Apply the Kinetic Molecular Theory to calculations involving specific heat.	Apply the Kinetic Molecular Theory to predict the behavior of matter.	Apply the Kinetic Molecular Theory to predict the behavior of matter.	Interpret factors that affect equilibrium and kinetics.	Apply the Kinetic Molecular Theory to predict the behavior of matter.	Apply the Kinetic Molecular Theory to predict the behavior of matter.	Total
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Student Performance by Question
Core 2 / S-23

Chemistry (Multiple-Choice ONLY)

Teacher: ____________________________

Analysis of Performance by Question of each student targeted for Remediation or Intervention
Enter a "1" in each box to indicate a missed question.

	Scientific Investigation										Atomic Structure and Periodic Relationships								Chemical Formulas and Reactions										
Student Name	Identify the correct equipment to collect the results of a safe controlled experiment.	Analyze results of a controlled experiment using significant digits conventions.	Recognize and follow safe and appropriate laboratory procedures.	Calculate the percent error in the results of a controlled experiment.	Identify proper laboratory safety techniques when handling chemicals.	Recognize and follow safe and appropriate laboratory procedures.	Identify conclusion of historical experiments.	Interpret results using appropriate error analysis.	Calculate percent error.	Apply an understanding of precision and accuracy.	Identify the general trends of the Periodic Table.	Understand the charge and mass of subatomic particles.	Recognize that periodic trends can be determined from the periodic table.	Identify characteristics of subatomic particles.	Use bonding principles to explain the interaction of substances during chemical changes.	Recognize that properties can be determined from the periodic table.	Apply information provided by the periodic table to make comparisons of matter.	Evaluate information to calculate properties of isotopes.	Determine the molecular formula of a compound, given certain information.	Classify chemical equations.	Interpret factors that affect equilibrium and kinetics.	Recognizes the function of a catalyst.	Describe compounds and bonding to provide appropriate chemical names.	Describe compounds and bonding to provide appropriate chemical names.	Use bonding principles to explain the interaction of substances during chemical changes.	Use bonding principles to explain the interaction of substances.	Recognize different types of reactions.	Interpret factors that affect equilibrium and kinetics.	Balance chemical equations.
Total	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0

					Molar Relationships								Phases of Matter and Kinetic Molecular Theory								
Recognize bond types.	Balance chemical equation and select specified coefficient.	Classify chemical equations.	Balance chemical equations.	Describe ionic bonding.	Interpret factors that affect solution chemistry.	Apply molar relation ships to perform calculations involving molar concentrations.	Interpret factors that affect solution chemistry.	Apply molar relation ships to perform calculations involving molar stiochiometry.	Calculate the molarity of a solution.	Apply molar relation ships to perform calculations involving molar concentrations.	Calcuates number of particles by using Avogadro's number when mass is given.	Calculate the volume of a gaseous product at STP given moles of reactant and a balanced chemical equation.	Calculate heat needed to raise temperature of a given amount of substance.	Understand the parameters affecting vapor pressure.	Predict the behavior of matter and interactions between particles in colligative properties.	Interpret factors that affect equilibrium and kinetics.	Apply the Kinetic Molecular Theory to predict the behavior of matter.	Apply molar relationships to perform calculations involving molar conversions.	Apply the Kinetic Molecular Theory to predict the behavior of matter.	Understand the relationship between kinetic energy and temperature.	Total
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Student Performance by Question
Core 3 / S-33

Chemistry (Multiple-Choice ONLY)

Teacher: _____

Analysis of Performance by Question of each student targeted for Remediation or Intervention
Enter a "1" in each box to indicate a missed question.

	Scientific Investigation										Atomic Structure and Periodic Relationships								Chemical Formulas and Reactions										
Student Name	Calculate percent error.	Determine the appropriate type of probe.	Demonstrate an understanding of proper laboratory techniques.	Apply an understanding of mathematical manipulations.	Identify proper laboratory safety techniques when handling chemicals.	Identify variables in a investigation.	Identify conclusion of historical experiments.	Calculate percent error.	Interprets solubility curves for various substances.	Apply an understanding of precision and accuracy.	Identify conclusion of historical experiments.	Identify the general trends of the Periodic Table.	Use a chart to determine an ion's charge.	Understand the charge and mass of subatomic particles.	Apply information provided by the periodic table to construct models.	Apply knowledge of isotopes.	Identify the similarity in the chemical properties of a given family.	Recognize and follow safe and appropriate laboratory procedures.	Determine the molecular formula of a compound, given certain information.	Recognizes the function of a catalyst.	Choose the chemical formula for a given compound.	Choose the chemical formula for a given compound.	Balance chemical equations.	Apply an understanding of chemical equilibrium in solutions.	Recognize an ionic bond.	Explain the function of a catalyst.	Recognize different types of reactions.	Apply an understanding of chemical nomenclature.	Examine a chemical equation to determine reaction type.
Total	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0

					Molar Relationships								Phases of Matter and Kinetic Molecular Theory								
Recognize bond types.	Recognize a synthesis reaction.	Balance chemical equation and select specified coefficient.	Apply an understnding of chemical nomenclature.	Determine the molecular formula of a compound, given certain information.	Apply molar relationships to perform calculations involving stoichiometry.	Determine the number of particles when moles of elements are given.	Determine the concentration of ions given pH or pOH.	Apply an understanding of molar volume in gases at STP.	Calculate the molarity of a solution.	Calculate quantity of compound produced from balanced equation.	Recognizes properties of electrolyte solutions.	Calculates number of particles by using Avogadro's number when mass is given.	Calculate heat needed to raise temperature of a given amount of substance.	Understand the parameters affecting vapor pressure.	Recognize the relationship between temperature and volume in gases.	Understand the effects of solutes on colligative properties.	Understand the effect of intermolecular forces on molar heat of vaporization.	Identify causes for boiling point evaluation and freezing point depression.	Understand the relationship between kinetic energy and temperature.	Calculate the partial pressure of oxygen in a mixture.	Total
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